

## The halogens - a summary

Halogen	Colour	State	Use
Fluorine, F <sub>2</sub>			
Chlorine, Cl <sub>2</sub>			
Bromine, Br <sub>2</sub>			
Iodine, I <sub>2</sub>			

**Colours:** Black, orange, green, yellow, grey, white

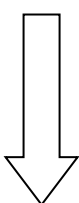
**States:** solid, liquid, gas

**Uses:** photography, toothpaste, antiseptic, bleach

Write an equation to show how:

- Lithium Chloride is formed from lithium and chlorine
  
- Sodium Iodide is formed

## Displacement Reactions - Checking understanding

<b>Halogen</b>		<b>Reactivity</b>
Fluorine		Most
Chlorine		
Bromine		
Iodine		Least



Chlorine **DISPLACES** bromine because chlorine is more reactive.

Use the reactivity series to help you decide whether displacement will happen in these cases. Complete the equations. If displacement does not occur write "no reaction".

Bromine + Potassium Chloride →

Bromine + Sodium Iodide →

Iodine + Potassium Chloride →

Iodine + Lithium Bromide →

Fluorine + Potassium Chloride →