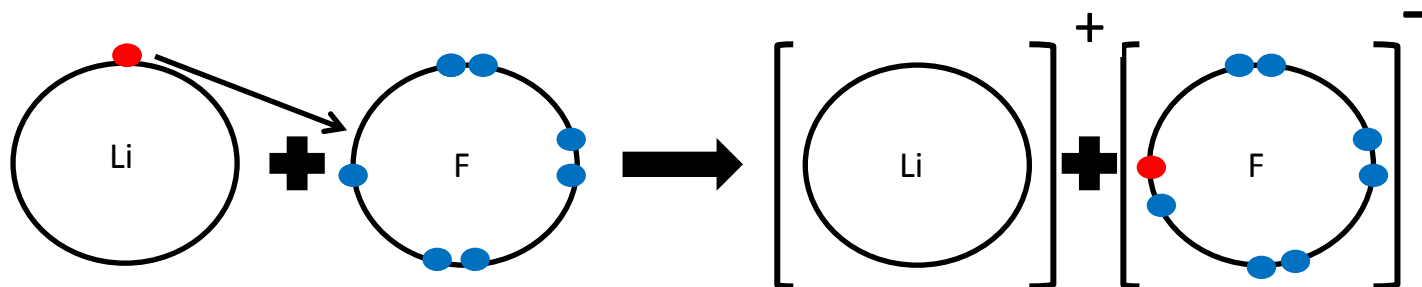


Ionic Bonding

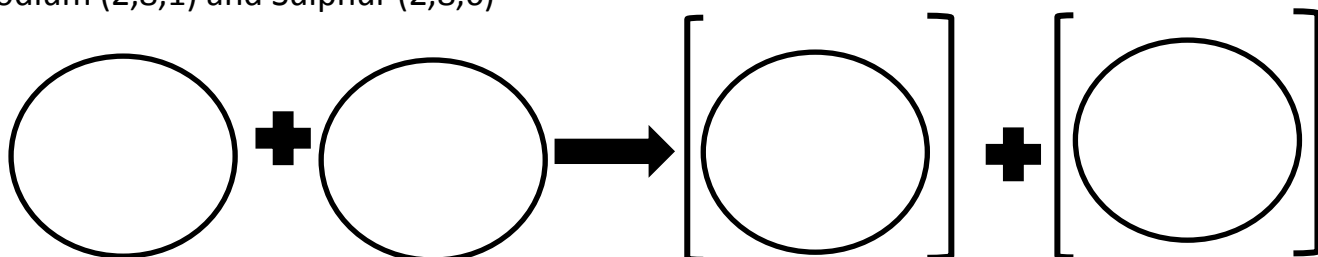
1. Decide how many electrons are in the outer shell of each atom and mark them on.
2. Show how the electrons are transferred.
3. Show how many electrons the ions now have.
4. Work out the charge on each ion.

Example Lithium (2,1) and Fluorine (2, 7)

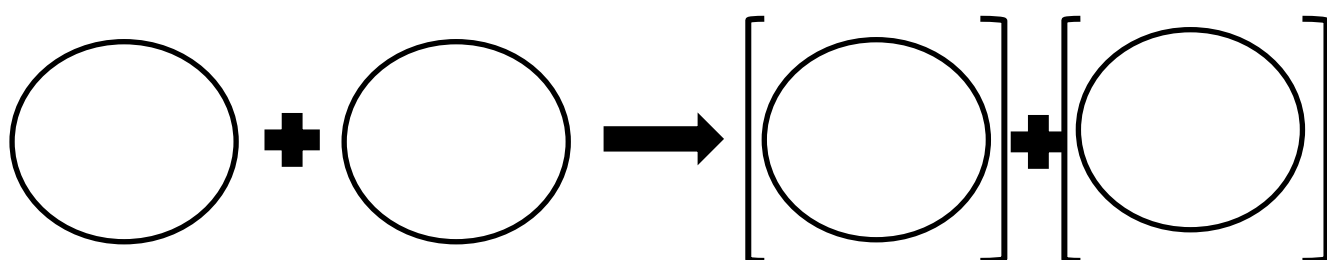


Try these ones yourself – you might need to draw more atoms and ions

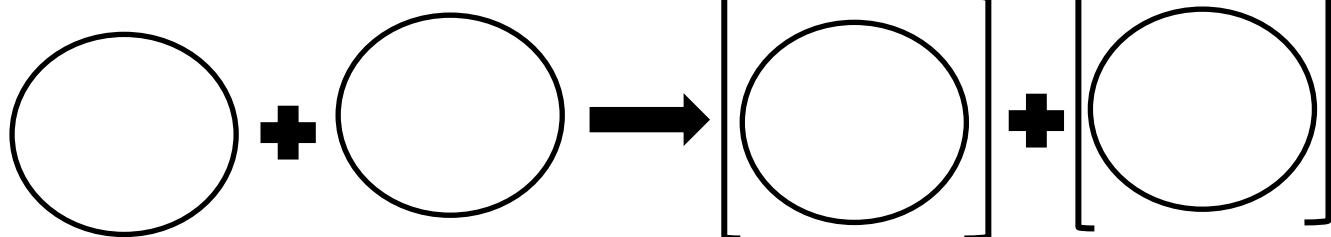
Sodium (2,8,1) and Sulphur (2,8,6)



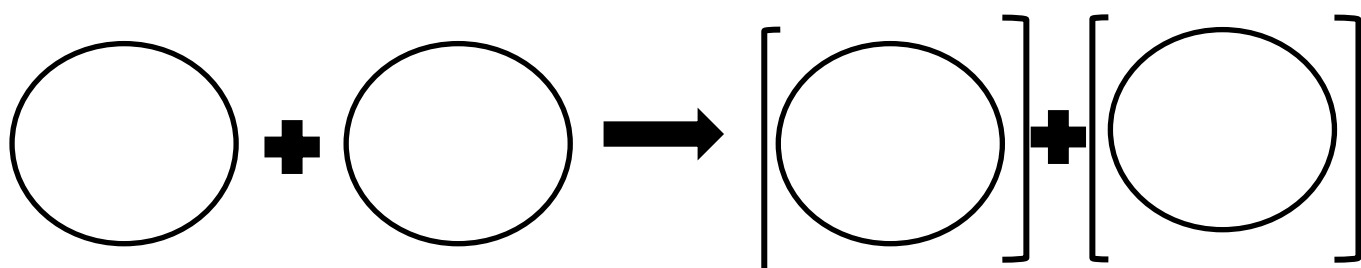
Magnesium (2,8,2) and Sulphur (2,8,6)



Magnesium (2,8,2) and Fluorine(2,7)



Aluminium (2,8,3) and Nitrogen (2,5)



Now try these – you will have to draw your own diagrams

1. Sodium (2,8,1) and Oxygen (2,6)
2. Calcium (2,8,8,2) and Chlorine (2,8,7)
3. Aluminium(2,8,3) and Fluorine (2,7)
4. Aluminium (2,8,3) and Oxygen(2,6)

With these you will have to work out the number of electrons each atom has as well.

5. Potassium and Sulphur
6. Boron and Phosphorous
7. Beryllium and Oxygen